AC ??/??/2014 Item no. ?.??

UNIVERSITY OF MUMBAI

Revised Syllabus
Program - **Bachelor of Engineering**
Course - **Chemical Engineering**
(Third year - Sem V and VI)

under
Faculty of Technology
(As per Credit Based Semester and Grading System from 2014-15)

University of Mumbai

Scheme for TE: Semester-V

Course Code	Course Name	Teaching Scheme			Credit Assigned			
		Theory	Pract.	Tut.	Theory	Pract.	Tut.	Total
CHC501	Chemical Engineering Thermodynamics - II	03	–	01	3.0	–	1.0	4.0
CHC502	Mass Transfer Operations - I (MTO-I)	03	–	01	3.0	–	1.0	4.0
CHC503	Heat Transfer Operations – I (HTO-I)	03	–	01	3.0	–	1.0	4.0
CHC504	Chemical Reaction Engineering - I (CRE-I)	03	–	01	3.0	–	1.0	4.0
CHC505	Chemical Technology	03	–	–	3.0	–	–	3.0
CHC506	Business Communication & Ethics	–	02* + 02	–	–	–	–	2.0
CHL507	Chemical Engg Lab (MTO-I)	–	03	–	–	1.5	–	1.5
CHL508	Chemical Engg Lab (CRE-I)	–	03	–	–	1.5	–	1.5
CHL509	Chemical Engg Lab (HTO-I)	–	03	–	–	1.5	–	1.5
CHL510	Chemical Engg Lab (Synthesis)	–	03	–	–	1.5	–	1.5
	Total	15	16	04	15.0	6.0	6.0	27.0

*Theory for entire class.

Examination Scheme

Course Code	Course Name	Examination Scheme							
		Theory marks				Term Work	Pract.	Oral	Total
		Internal Assessment			End Sem. Exam				
		Test 1	Test 2	Avg. of Test 1 and Test 2					
CHC501	Chemical Engineering Thermodynamics - II	20	20	20	80	25	–	–	125
CHC502	Mass Transfer Operations - I (MTO-I)	20	20	20	80	25	–	–	125
CHC503	Heat Transfer Operations – I (HTO-I)	20	20	20	80	25	–	–	125
CHC504	Chemical Reaction Engineering - I (CRE-I)	20	20	20	80	25	–	–	125
CHC505	Chemical Technology	20	20	20	80	–	–	–	100
CHC506	Business Communication & Ethics	–	–	–	–	50	–	–	50
CHL507	Chemical Engg Lab (MTO-I)	–	–	–	–	–	25	–	25
CHL508	Chemical Engg Lab (CRE-I)	–	–	–	–	–	25	–	25
CHL509	Chemical Engg Lab (HTO-I)	–	–	–	–	–	25	–	25
CHL510	Chemical Engg Lab (Synthesis)	–	–	–	–	–	–	25	25
	Total			100	400	100	75	75	750

Course Code	Course Name	Credits
CHC501	Chemical Engineering Thermodynamics - II	4.0

Prerequisites

Chemical Engineering Thermodynamics – I, Engineering Mathematics.

Course Objectives

The course objectives are

- The student should be able to relate thermodynamics to the Chemical Engineering Problems.
- The students should be able to use thermodynamics rules to find the equilibrium in phases.
- The students should be able to calculate and trace the equilibrium concentration and conversions of a reversible reaction.
- The students should be able to calculate the actual power required for given duty of refrigeration.

Course Outcomes

The student learn the application of First law and second law to the problem of phase equilibrium and reaction equilibrium . The students also learn to calculate the refrigerant flow rate for a given duty of refrigeration. This helps in estimating the compressor sizes and loads for refrigeration. The calculation of phase equilibria and the understanding of it is a fundamental concept to design of mass transfer equipment.

Detail syllabus

Module	Contents	No. of hrs
1	**Reaction Thermodynamics:** Calculation of heat of reaction for batch reactors, Calculation of heat of reaction for continuous reactors.	05
2	**Fundamentals of Phase Equilibria:** Concept of equilibrium in phases, The theory of ideal and non ideal solutions, Thermodynamic equations of Vapor Liquid Equilibrium for ideal and non ideal solutions, Liquid Liquid and Solid Liquid equilibria.	12
3	**Reaction Equilibria:** Representation of reaction stoichiometry, Concept of reaction equilibria, single and multiple reactions, Degrees of freedom for single and multiple reactions.	10
4	**Refrigeration:** Theory of refrigeration, Vapor Absorption Refrigeration, Vapor Absorption Refrigeration, Estimation of refrigerant flow rate and power of compression.	07
5	**Methods for estimation of Thermodynamics properties:** Estimation methods for critical parameters, Estimation method for Mixture Enthalpy and Entropy.	05

References

1. Stanley I. Sandler, Chemical, Biochemical, and Engineering Thermodynamics, 4 ed., Wiley Student Edition

2. M.J. Moran, H.N. Shapiro, Fundamentals of Engineering Thermodynamics, 6 ed., Wiley Student Edition

3. Peter Atkins, Physical Chemistry, 9 ed., Oxford University Press.

Note for the teacher/instructors: The teachers should encourage the student to use computer for solving problems. It would be worth mentioning that Microsoft Excel suffices for solving most of the problems in the syllabus. A total of twelve assignments and tutorials together should be given to the students at regular intervals. Students should be encouraged to submit assignment using word processor and as far as possible they should be allowed to submit it online in some form. As far as possible it should be multiple choice questions for problem based in mid term tests.

Course Code	Course Name	Credits
CHC502	Mass Transfer Operations - I (MTO-I)	4.0

Prerequisites

Knowledge of chemistry, physics, physical chemistry, mathematics, process calculations and unit operations.

Course Objectives

To give insight of mass transfer basic principle and mass transfer mechanisms.

Course Outcomes

At the end of the course students will be able to ...

- demonstrate the knowledge of mass transfer by applying principles of diffusion, mass transfer coefficients, and interphase mass transfer.
- understand the concept and operation of various types of gas-liquid contacts equipments.
- determine NTU, HTU, HETP and height of packed bed used for Absorption and Humidification operations.
- find time required for drying.

Detail syllabus

Module	Contents	No. of hrs
1	**Molecular Diffusion in Gases and Liquid:** Basics of Molecular Diffusion, Fick's First Law of Molecular Diffusion, Various fluxes and relations between them, Molecular Diffusion in binary gas mixtures – Steady state diffusion of one component in non-diffusing second component, Equimolal counter diffusion of two components. Molecular Diffusion in binary liquid solutions – Steady state diffusion of one component in non-diffusing second component, Steady State Equimolal counter diffusion of two components. Diffusivity of gases. Theoretical and experimental determination of diffusivities, Diffusivities of liquids – Theoretical Determination. Diffusion in Solids: Ficks law of diffusion in solids, Types of Solid Diffusion, Diffusion through Polymers, Diffusion through Crystalline Solids, Diffusion in Porous Solids	08
2	**Mass Transfer Coefficients:** Definition of Mass Transfer Coefficient, F-Type and K-Type Mass Transfer Coefficients and relations between them, Mass Transfer Coefficients in Laminar and Turbulent Flow. Heat, Mass and Momentum Transfer Analogies and dimensionless numbers, Interphase Mass Transfer – Individual and Overall Mass Transfer Coefficients and relation between them. Methods of contacting two insoluble phases – Continuous Contact, Stage-wise Contact.	08

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Module	Contents	No. of hrs
3	**Equipments for Gas-Liquid Contacting:** Classification of equipments for gas-liquid contacting • Gas dispersed and liquid continuous phase – Sparged Vessels (Bubble Columns), Mechanically Agitated Vessels, Tray Towers. • Liquid dispersed phase and gas continuous phase – Venturi Scrubbers, Wetted Wall Towers, Spray Towers and Spray Chambers, Packed Towers. Comparison of Packed Towers with Tray Towers.	06
4	**Gas Absorption:** Solubility of gases in liquids, Effect of temperature and pressure on solubility, Ideal and Non-ideal solutions, Choice of solvent for gas absorption, Single component gas absorption – Cross Current, Co-current, Countercurrent, Multistage Counter current Operation. Absorption with Chemical Reactions.	06
5	**Drying:** Introduction to drying, Equilibrium, Different types of moisture contents, Rate of Drying and drying curve, Batch Drying and calculation of time of drying, Continuous	06
6	**Humidification and Dehumidification:** Introduction, Vapour Pressure Curve, Properties of Vapour-Gas mixtures [Understanding various terms], Theory of wet bulb temperature, Adiabatic Saturation Curves,Humidity Charts, Adiabatic operation : (Air water systems) water coolers, cooling towers	06

References

1. Treybal R.E. , Mass transfer operation, 3 Ed., McGraw Hill New York, 1980.
2. McCabe W.L. and Smith J.C., Unit operation in chemical engineering, 5 Ed., McGraw Hill New York 1993.
3. Geankoplis C.J., Transport processed and unit operations, Prentice Hall , New Delhi 1997.
4. Coulson J.M. Richardson J.F., Backhurst J.R. and Harker J.H., Coulson and Richardson chemical engineering, vol 1 & 2, Butterworth Heinman, New Delhi, 2000.
5. R.K.Sinnot (Ed) Coulson and Richardson chemical engineering, vol 6, Butterworth Heinman, New Delhi, 2000.

Course Code	Course Name	Credits
CHC503	Heat Transfer Operations – I (HTO-I)	4.0

Prerequisites

Laws of thermodynamics, Units and dimensions, Fluid flow principles, Solution techniques of ordinary and partial differential equations.

Course Objectives

- Students should be able to calculate rate of heat transfer by all three modes of heat transfer.
- Understand the basic principles involved in mechanism and calculation of heat transfer rates.
- Able to deal with most common types of unsteady state operations of heat transfer.
- Should become familiar with equipments, used for heat transfer in industry.

Course Outcomes

Upon completion of this course the learners will be acquainted to process design concept of heat transfer equipments and prepared for heat transfer equipment design study.

Detail syllabus

Module	Contents	No. of hrs
1	**Introduction:** fundamentals of heat transfer, basic modes of heat transfer. Concept of driving force and heat transfer coefficients, rate expressions for three modes i. e. conduction, convection, radiation	02
2	**Steady state conduction:** Fourier's Law, thermal conductivity, conduction through a flat slab, composite slab, conduction through a cylinder, composite cylinder, conduction through sphere, composite sphere. Critical radius of insulation. Concept of thermal resistance, fouling factors, Wilson plot, calculation of overall heat transfer coefficients.	05
3	**Unsteady state conduction:** Lumped Parameter Analysis -systems with negligible internal resistance. Biot number, Fourier number, Heating a body under conditions of negligible surface resistance,, heating a body with finite surface and internal resistance, heat transfer to a semi-infinite wall.	04
4	**Heat transfer by convection:** Fundamental considerations in convective heat transfer, significant parameters in convective heat transfer such as momentum diffusivity, thermal diffusivity, Prandtl number, Nusselt number, dimensional analysis of convective heat transfer-Natural and Forced convection, convective heat transfer correlations for internal and external flows, equivalent diameter for heat transfer, estimation of wall temperature, correlations for heat transfer by natural convection from hot surfaces of different geometries and inclination.	07

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Module	Contents	No. of hrs
5	**Heat transfer in condensation and boiling:** Introduction, types of condensation, Nusselt's theory of condensation, correlations for vertical and horizontal tube, plate, for stack of tubes etc. Heat transfer to boiling liquids, regimes of pool boiling of saturated liquid, correlations for estimating the boiling heat transfer coefficients.	05
6	**Steam:** Properties of steam. Steam generation by utilizing process waste heat, efficient use of steam in plant.	04
7	**Heat transfer through extended surfaces:** longitudinal, transverse and radial fins, calculations with different boundary conditions, efficiency and effectiveness of fin, calculation of rate of heat transfer.	03
8	**Heat Exchangers:** Classification and types of heat exchangers, Double pipe heat exchanger, calculation of LMTD, effectiveness NTU method. Introduction to Shell and Tube Heat Exchanger. heat transfer in agitated vessel	05
9	**Radiation heat transfer:** Emissivity, absorptivity, black body, grey body, opaque body, Stephan Boltzmann law, Kirchoff's law. Equations for rate of heat transfer by radiation for various cases. Basic unsteady state radiation heat transfer.	04

References

1. D. Q. Kern, Process Heat Transfer, McGraw Hill, 1997.
2. Incropera Frank P., Dewitt David P., Bergman T. L., Lavine A. S., Seetharamu K. N., Seetharam T. R., Fundamentals of Heat and Mass Transfer, Wiley, 2014.
3. Holman, J. P., Heat Transfer, 9 ed., McGraw Hill, 2008.
4. R. K. Sinnott, Coulson & Richordsons Chemical Engineering Design, Vol. 6, Elsevier Butterworth-Heinemann.
5. J. M. Coulson and J. F. Richardson with J. R. Backhurst and J. H. Harker, Coulson & Richordsons Chemical Engineering Design, Vol. 1 & 2, Elsevier Butterworth-Heinemann, 1996.
6. W. D. Seider, J. D. Seader, D. R. Lewin, Product & Process Design Principles Synthesis, Analysis and Evaluation, John Wiley and Sons, Inc.
7. Robert W. Serth, Process Heat Transfer: Principles and Applications, Elsevier Science & Technology Books.
8. John H. Lienhard IV, John H. Lienhard V, A Heat Transfer Textbook, Phlogiston Press.
9. McCabe W.L., Smith J.C., Harriot P., Unit Oerations of Chemical Engineering, 5th ed., McGraw Hill, 1993

Course Code	Course Name	Credits
CHC504	Chemical Reaction Engineering - I (CRE-I)	4.0

Prerequisites

Students should know basic Chemistry pertaining to Chemical Reactions, Chemical formula etc. They are required to be aware of Chemical processes and unit operations used for the manufacturing of chemical products. Simple to complex numerical methods of solving one and two dimensional Mathematical equations.

Course Objectives

- Development of Kinetic model for homogeneous reactions giving emphasis on various types of reactions like reversible, irreversible, 1st order, 2nd order reactions, series parallel reactions, homogeneous catalytic reactions, autocatalytic reactions, reactions in adiabatic or non isothermal conditions.
- Development of design strategy for homogeneous reactions considering different types of reactors for example batch reactors, flow reactors, semi batch reactors, recycle reactors etc.Reactor design for reactions operating under adiabatic or non-isothermal conditions.

Course Outcomes

Students will be able to apply the knowledge they have gained to find the model equation and use this model to design the reactors used for homogeneous reactions taking place in isothermal or non isothermal conditions.

Detail syllabus

Module	Contents	No. of hrs
1	**Introduction and reaction kinetics of homogeneous systems:** Various types of reactions: Reversible Vs irreversible reactions. Homogeneous Vs heterogeneous reactions. Catalytic Vs non-catalytic reactions. Single vs multiple reactions. Auto catalytic reactions, Rate of reaction, Rate constants, Order/ molecularity. Formulation and solution of rate equations for batch reactors for simple and complex reactions. Effect of thermodynamic equilibrium. Temperature dependency-Various Theories. Reaction mechanism and it influence on kinetics, search for plausible mechanism via reaction kinetics	09
2	**Methods of analysis of experimental data:** For Constant volume & variable volume batch reactor – Integral method of analysis of experimental data, Differential method of analysis. Concept of half-life /fractional life. Over all order of irreversible reactions (initial rate method). Empirical rate equation for n^{th} order reactions. Analysis of complete rate of reactions. Partial analysis of rate of reaction. Reversible and irreversible reactions in parallel Reversible and irreversible reaction in series. Homogeneous catalysed reactions. Auto Catalytic reactions. Shifting order reactions	09

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Module	Contents	No. of hrs
3	**Design of reactor:** Ideal batch reactor and concept of batch time. Flow reactor and concept of space time / space velocity and holding time / residence time. Ideal mixed flow reactor (MFR) and plug flow reactor (PFR). Design for single reactions: Single reactor performance of reversible and irreversible first order, pseudo first order, second order reactions for MFR, PFR. Graphical and analytical techniques. Combination of reactors PFR in series / parallel, unequal size MFR in series, performance of the above for the first order and second order reactions. Recycle reactor and auto catalytic reactor. Semi batch reactor and recycle reactor. Design for complex reactions: Irreversible and Reversible reactions in series and parallel with same or different order in various combinations.	12
4	**Heat and pressure effects:** Heat of reaction and its variation with temperature. Variation of equilibrium constant and equilibrium conversion with temperature. Effect of temperature on reactor performance for adiabatic and non adiabatic operations. Case of exothermic reactions in mixed reactor. Optimum temperature progression. Multiple reactions- effect on product distribution. Temperature and scale effect on productivity of reactor. Various problems based on design of non-isothermal reactor are to be solved by using various numerical methods.	09

References

1. Levenspiel, O.,Chemical Reaction Engineering, 3 ed., John Wiley & Co.
2. Smith J.M., Chemical Engineering Kinetics, McGraw Hill.
3. Laidler, K.J., Chemical Kinetics, Tata McGraw Hill, 1997.
4. Hill C.G., Chemical Reaction Engineering.
5. Walas, Reaction Kinetics for Chemical Engineers, McGraw Hill.
6. Sharma M.M & L.K Doraiswamy, Heterogeneous Reactions, Vol 1
7. Fogler, H.S., Elements of Chemical Reaction Engineering, 4 ed.,PHI, 2008.

Course Code	Course Name	Credits
CHC505	Chemical Technology	3.0

Prerequisites

Knowledge of chemistry, physics, physical chemistry and mathematics. Knowledge of Unit Operations and Unit Processes. Knowledge of material balance and energy balance

Course Objectives

- To give students insight of different chemical processes.
- To understand development of process from its chemistry.
- To understand different engineering problems in process industry.

Course Outcomes

At the end of the course student will be able to :

- demonstrate various manufacturing processes,
- explain industrial processing and overall performance of any chemical process,
- find out the overall process aspects including yield, waste etc.,
- draw and illustrate the process flow diagram.

Detail syllabus

Module	Contents	No. of hrs
1	Introduction: Unit Operations and Processes Concept Used in Chemical Industries. General principles applied in studying an industry, phases of development of chemical industries in India. An overview on industries such as: vegetable oils & animal fats, natural waxes / resins, essential oils & Flavour ingredients Industry, Food & Agro-Products An overview of other industrially important products: Paints, Varnishes & lacquers, Soaps & Detergents, Dyes & Intermediates, Agrochemicals, Pharmaceuticals: Penicillin.	07
2	Manufacturing of Acids: Sulphuric Acid (DCDA Process), Nitric Acid, Acetic Acid & Phosphoric Acid (WET Process), Manufacturing of Fertilizers: Ammonia, Urea, Superphosphate (SSP, TSP) & Ammonium Sulphate	08
3	Sugar, starch & alcohol industries. Introduction to biodiesel processing. Chloro-Alkali Industries: Manufacturing of Caustic Soda, Hydrochloric Acid and Hydrogen, Soda Ash (Solvay and Dual Process).	07
4	Synthesis of Important Heavy Organic Chemicals and Intermediates : Styrene , Phenol, Purified Terephthalic acid.	07
5	Synthesis of Polymers: Polyethylene: LDPE, LLDPE and HDPE; Polyster Fibre, Nylon and PVC.	06

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Module	Contents	No. of hrs
6	Basic Building Blocks of Petrochemical Industry: Treatment of Crude oils and the products there from; refining vs. cracking; manufacture of Acetylene, Ethylene, Benzene Toluene, Xylene.	05

References

1. Austin, G. T., Shreve's Chemical Process Industries, 5 Ed., McGraw Hill International Edition.
2. Pandey, G. N., A text book of Chemical Technology, Vol. I and II., Vikas Publications, 1984
3. Rao, G. N. and Sittig, M. Drydens outlines of Chemical Technology for 21st Century, East West Press, 3rd edition
4. Heaton, C. A., An introduction to industrial chemistry, Leonard Hill, 1984
5. Thomson, R., Moden inorganic chemicals industries, Royal Society of chemistry, 2nd ed., 1994
6. Kirk-Other's, Encyclopedia of chemical technology, John Wiley and sons Inc., 4th ed. 1990
7. Ullmanns Encyclopedia of Industrial Chemistry, VCH, 1985
8. McKettas Encyclopedia of chemical processing and design, Marcel Dekker, 1999
9. Pletcher, D. and Walsh, F. C., Industrial Electro-chemistry, Chapman & Hall, 1990
10. Alok Adholeya and Pradeepkumar Dadhich, Production and Technology of Biodiesel: seeding a change, TERI Publication, New Delhi, 2008
11. NIIR Board of consultants and Engineers, The complete book on Jatropha (Biodiesel) with ash-wagandha, stevia, brahmi and Jatamansi Herbs (cultivation, processing and uses), Asia Pacific Business Press Inc.

Course Code	Course Name	Credits
CHC506	Business Communication & Ethics	2.0

Course Objectives

- To inculcate in students professional and ethical attitude, effective communication skills, teamwork, skills, multidisciplinary approach and an ability to understand engineers social responsibilities.
- To provide students with an academic environment where they will be aware of the excellence, leadership and lifelong learning needed for a successful professional career.
- To inculcate professional ethics and codes of professional practice.
- To prepare students for successful careers that meets the global Industrial and Corporate requirement provide an environment for students to work on Multidisciplinary projects as part of different teams to enhance their team building capabilities like leadership, motivation, teamwork etc.

Course Outcomes

A learner will be able to

- Communicate effectively in both verbal and written form and demonstrate knowledge of professional and ethical responsibilities,
- participate and succeed in Campus placements and competitive examinations like GATE, CET,
- possess entrepreneurial approach and ability for life-long learning,
- have education necessary for understanding the impact of engineering solutions on Society and demonstrate awareness of contemporary issues.

Detail syllabus

Module	Contents	No. of hrs
1	Report Writing: Objectives of report writing Language and Style in a report Types of reports Formats of reports: Memo, letter, project and survey based	7
2	Technical Proposals Objective of technical proposals Parts of proposal	2
3	Introduction to Interpersonal Skills Emotional Intelligence Leadership Team Buliding Assertiveness Conflict Resolution Negotiation Skills Motivation Time Management	7
4	Meetings and Documentation Strategies for conducting effective meetings Notice Agenda Minutes of the meeting	2
5	Introduction to Corporate Ethics and etiquettes Business Meeting etiquettes, Interview etiquettes, Professional and work etiquettes, Social skills Greetings and Art of Conversation Dressing and Grooming Dinning etiquette Ethical codes of conduct in business and corporate activities (Personal ethics, conflicting values, choosing a moral response, the process of making ethical decisions)	2
6	Employment Skills Cover letter Resume Group Discussion Presentation Skills Interview Skills	6

References

1. Fred Luthans, Organizational Behavior , Mc Graw Hill, edition
2. Lesiker and Petit, Report Writing for Business , Mc Graw Hill, edition
3. Huckin and Olsen, Technical Writing and Professional Communication, McGraw Hill
4. Wallace and Masters, Personal Development for Life and Work , Thomson Learning, 12th edition
5. Heta Murphy, Effective Business Communication , Mc Graw Hill, edition
6. R.C Sharma and Krishna Mohan, Business Correspondence and Report Writing,
7. B N Ghosh, Managing Soft Skills for Personality Development, Tata McGraw Hill. Lehman,
8. Dufrene, Sinha, BCOM, Cengage Learning, 2nd edition
9. Bell . Smith, Management Communication Wiley India Edition,3rd edition.
10. Dr. K. Alex ,Soft Skills, S Chand and Company
11. Dr.KAlex,SoftSkills,S Chand and Company
12. R.Subramaniam, Professional Ethics Oxford University Press 2013.

Course Code	Course Name	Credits
CHL507	Chemical Engg Lab (MTO-I)	1.5

Concept for experiments

The laboratory work shall consist of a record of minimum eight experiments performed during the term. The design of experiments should cover all concepts (such as Mass transfer coefficient, Gas liquid contacts, Absorption, Drying, Humidification etc.) mentioned in the syllabus. Each and every experiment should conclusively demonstrate / verify the theory. The students should be able to explain variations (if any) between observed and expected results based on technical knowledge. Each experimental report should contain a discussion of the results obtained.

Course Code	Course Name	Credits
CHL508	Chemical Engg Lab (CRE-I)	1.5

Concept for experiments

Minimum 8 experiments need to be performed by the students on following concepts.

- Effect of concentration and temperature on reaction rate.
- Batch reactor.
- Arrhenius constants.
- Differential and integral analysis.
- Acidic hydrolysis.
- Condensation polymerisation kinetics.
- Constant flow stirred tank reactor (CSTR).
- Plug flow reactor (PFR).
- CSTRs connected in series.
- PFR-CSTR combination in series.

Course Code	Course Name	Credits
CHL509	Chemical Engg Lab (HTO-I)	1.5

Concept for experiments

Minimum seven practical including experiments on conduction, unsteady state conduction, forced and natural convection, condensation, heat exchangers should be done. These can include any additional experiment based on the syllabus.

Course Code	Course Name	Credits
CHL510	Chemical Engg Lab (Synthesis)	1.5

Concept for experiments

Concept for experiments to be designed by instructor is students should developed an approach towards engineering a chemical process. Following are some of the suggested processes,

- Preparation of a soap.
- Preparation of a detergent.
- Preparation of paper.
- Preparation of polymer product.
- Preparation of a pharmaceutical product.
- Preparation of a membrane.
- Preparation of a nano-particles.
- Preparation of a dye.
- Preparation of rubber.
- Preparation of a biochemical.
- Preparation of biodiesel.
- Preparation of a food product.
- Hydrogenation of oil.

examples of few lab prepared chemicals along with raw materials can be

Sr. No.	PREPARETION	Chemicals required	Apparatus/ glass-ware required
1	SOAP	Sodium hydroxide (20% solution), ethanol saturated solution of sodium chloride ,calcium chloride (5% solution), magnesium chloride (5% solution), ferric chloride (5% solution), cooking oil, phenolphthalein indicator solution.	250-mL beaker, 100- mL beaker; wire gauze; laboratory burner; glass stirring rod; test tubes; filter flask and Büchner funnel; filter paper ;graduated cylinder
2	ALUM FROM ALUMINUM	Aluminum can or aluminum metal, Crushed ice, 9M H_2SO_4, 1.5M KOH solution, Methanol, $NaHCO_3$ (sodium bicarbonate)	Glass filter funnel, Büchner filter funnel, filter paper, Steel wool, two 150 mL and two 150 ml beakers, 500 ml beaker, thermometer, ruler, stirring rod.

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Sr. No.	PREPARETION	Chemicals required	Apparatus/ glassware required
3	ASPRIN	2 gm salicylic acid, 5.0 ml of acetic anhydride, five drops of 85% phosphoric acid, distilled water	burette clamp, burner, stand with iron ring, wire gauze, ice bath,50 ml flask beaker, Büchner funnel aspirator
4	METHYL ORANGE	0.29 g of anhydrous sodium carbonate, 1.0 g of sulfanilic acid monohydrate, 0.375 g of sodium nitrite, 0.7 ml of dimethylaniline and 0.5 mL of glacial acetic acid, 10% aqueous sodium hydroxide, 1.25 ml of concentrated hydrochloric acid	50 ml Erlenmeyer flask, filter,100 ml beaker, test tube
5	THIOKOL RUBBER	Sodium hydroxide solution, 1M Sulfur 1,2-dichloroethane distilled or deionized water	Copper wire, approximately 6 inches long (15 cm); two 10 ml vials with teflon cap liners, two 400 ml beakers ,10 ml graduated cylinder ,glass pipette (dropper), hot plate, chemical resistant gloves
6	RUBBER BALL FROM RUBBER LATEX	15 ml rubber latex, 15 ml vinegar, 15 ml water	Two paper cups (5 ounce), stirring rod (popsicle stick or equivalent), small bucket or large beaker (1000 ml or larger)
7	p-BROMO-NITROBENEZENE FROM BROMOBENEZENE	Conc. H_2SO_4, conc. HNO_3, bromobenzene, ethyl alcohol, conical flask, funnel, filter paper, water Bath.	Conical flask,funnel, filter paper, water bath.
8	DETERGENT	Dodecanol (dodecyl alcohol), sulphuric acid, concentrated sodium hydroxide, 6M phenolphthalein solution, 1% sodium chloride	Erlenmeyer flask, 125 ml beakers, 400 ml, 150 ml, 100 ml graduated cylinders, 10 ml, 25 ml, 125 ml funnel, spatula, stirring rod, Cheese cloth, watch glass, scissors